COURSE DATA

Data Subject	
Code	34186
Name	Chemistry laboratory II
Cycle	Grade
ECTS Credits	6.0
Academic year	2018 - 2019

Study (s)

Degree	Center	Acad. year	Period
1110 - Degree in Chemistry	Faculty of Chemistry	1	Second term

Subject-matter

Degree	Subject-matter	Character
1110 - Degree in Chemistry	1 - Chemistry	Basic Training

Coordination

Name	Department
MILIAN MEDINA, BEGOÑA	315 - Physical Chemistry

SUMMARY

This subject is compulsory of basic character that is taught in the second semester of the first year of the Degree in Chemistry, with a volume of 6 credits. Together with the "Chemistry Lab I" (also compulsory of basic character, but taught in the first semester), it is intended, essentially, that the student learn the operation of a chemical laboratory, as well as the basic techniques of work that will develop in it. In this way, the essential foundations will be established so that the experiences of the different branches that make up the discipline can subsequently be successfully addressed.In this specific subject the bases developed in the previous laboratory will be consolidated: the security, analysis and interpretation of data necessary for the development of any chemical experience, as well as the management and data processing that take place in any chemical laboratory. For this purpose, experiments will be carried out in which different basic techniques must be applied, applying them to more elaborate experiments. Experiments will be carried out on kinetics and thermodynamics of chemical reactions, equilibria and electrochemistry.It is required that the students already have consolidated some knowledge about safety and laboratory management, waste discrimination, preparation of memories and laboratory notebooks, correct use of material and products, data processing and realization of basic techniques developed in the Laboratory of Chemistry I.

PREVIOUS KNOWLEDGE

Relationship to other subjects of the same degree

There are no specified enrollment restrictions with other subjects of the curriculum.

Other requirements

Knowledge about laboratory safety and management, waste discrimination, preparation of reports and laboratory notebooks, correct use of materials and products, data processing and basic techniques developed in the Chemistry Laboratory I. In addition, it is assumed that the students know and use, in a basic but clear way, the concepts taught in the last year of High School Chemistry.

COMPETENCES (RD 1393/2007) // LEARNING OUTCOMES (RD 822/2021)

1108 - Degree in Chemistry

- Develop capacity for analysis, synthesis and critical thinking.
- Show inductive and deductive reasoning ability.
- Demonstrate leadership and management skills, entrepreneurship, initiative, creativity, organization, planning, control, leadership, decision making and negotiation.
- Demonstrate ability to work in teams both in interdisciplinary teams and in an international context.
- Demonstrate ability to communicate information, ideas, problems and solutions to both specialist and non-specialist audiences and using information technology, as appropriate.
- Demonstrate a commitment to ethics, equality values and social responsibility as a citizen and as a professional.
- Acquire a permanent sensitivity to quality, the environment, sustainable development and the prevention of occupational hazards.
- Demonstrate knowledge of the main aspects of chemical terminology, nomenclature, conventions and units.
- Demonstrate knowledge of the main types of chemical reaction and their main characteristics.
- Demonstrate knowledge of the principles of thermodynamics and kinetics and their applications in chemistry.
- Demonstrate knowledge and understanding of essential facts, concepts, principles and theories related to the areas of chemistry.
- Solve qualitative and quantitative problems following previously developed models.
- Evaluate, interpret and synthesise chemical data and information.
- Handle chemicals safely.

- Carry out standard experimental procedures involved in synthetic and analytical work, in relation to organic and inorganic systems.
- Interpret data from observations and measurements in the laboratory in terms of their significance and the theories that underpin them.
- Evaluate the risks in the use of chemicals and laboratory procedures.
- Recognise and evaluate chemical processes in daily life.
- Understand the qualitative and quantitative aspects of chemical problems.
- Students must have acquired knowledge and understanding in a specific field of study, on the basis of general secondary education and at a level that includes mainly knowledge drawn from advanced textbooks, but also some cutting-edge knowledge in their field of study.
- Students must have the ability to gather and interpret relevant data (usually in their field of study) to make judgements that take relevant social, scientific or ethical issues into consideration.
- Students must be able to communicate information, ideas, problems and solutions to both expert and lay audiences.
- Express oneself correctly, both orally and in writing, in any of the official languages of the Valencian Community.
- Have basic skills in the use of information and communication technology and properly manage the information obtained.

LEARNING OUTCOMES (RD 1393/2007) // NO CONTENT (RD 822/2021)

At the end of the module "Laboratory of Chemistry II", the student must be able to:

- Distinguish the types of cohesion forces present in chemical compounds.
- Predict and check the variation of the melting points according to the type of cohesion force.
- Predict and interpret the variation of the solubility of several chemical compounds and relate it to the type of cohesive force.
- Solubilize a carboxylic acid or an amine that are poorly soluble in water.
- Prepare a reflux system.
- Properly dose the reagents to carry out a synthesis of a simple chemical.
- Isolate by vacuum filtration the product of an organic synthesis reaction.
- Purify a product obtained in a reaction, by means of crystallization.
- Determine the purity of a solid product by determining the melting point and using the thin layer chromatography technique.
- Determine if a compound obtained by organic synthesis is more or less basic through the use of a suitable reagent and indicator.
- Use a calorimeter and perform the appropriate test to determine its heat capacity.
- Determine the enthalpy variation associated with the neutralization of a strong base with a strong acid.
- Determine the enthalpy variation associated with the dissolution of an ionic solid.
- Analyze the influence of temperature on the solubility of an ionic solid.
- Analyze the effect of the concentration of H^+ ions on the chromate-dichromate equilibrium.
- Deduce and check the effect of the common ion in equilibria in simple aqueous solutions (acetic

acid, ammonia, etc.)
- Analyze the redissolution of precipitates of metal hydroxides by the effect of several factors (addition of an acid, formation of a complex, etc.).
- Confirm the existence of reversible and irreversible reactions.
- Analyze the influence of temperature on complex ion equilibria (e.g. of Co^{2+}).
- List the main colligative properties.
- Express the concentration of a solution in terms of molality.
- Prepare a refrigeration mixture that reaches around -12 ºC.
- Construct a cooling curve of a pure solvent and a solution of a non-electrolyte compound (and deduce the melting point from it).
- Calculate the molar mass of a test compound (non-electrolyte) from the measurement of the cryoscopic depression.
- Experimentally determine the rate constant and the order of a reaction using a photocolorimetric technique.
- Use the UV-visible spectrophotometer to experimentally measure the variation of the absorbance of the discoloration reaction of the violet crystal (VC) in basic medium.
- Get the graph of the variation of the concentration versus time in the kinetics of discoloration of the VC.
- Calculate the partial reaction orders and the absolute rate constant in the discoloration reaction of the VC.
- Analyze the dependence of the decomposition rate of hydrogen peroxide catalyzed with iodide ion according to three variables: the H_2O_2 concentration, the catalyst concentration and the temperature.
- Follow the evolution of the decomposition reaction of H_2O_2 with time from measurements of the volume of oxygen evolved.
- Estimate the initial rate of the decomposition reaction of H_2O_2 for different concentrations of H_2O_2, of catalyst and at different temperatures.
- Express the rate law (reaction order and rate constant) of the decomposition kinetics of H_2O_2.
- Calculate the activation energy of the decomposition reaction of H_2O_2.
- Prepare the equipment to perform a potentiometric titration (burette, pH-meter, etc.).
- Perform the standardization of a NaOH solution, using a suitable primary standard.
- Obtain the data and represent the titration curves of a strong acid and a weak acid with the standardized solution of NaOH, and calculate the exact concentrations of both acid solutions.
- Determine the water self-ionization constant from the strong acid-base strong titration curve.
- Determine the acidity constant of acetic acid from the weak acid-base titration curve.
- Prepare different types of buffer solutions.
- Analyze the effect of adding bases or acids to buffer solutions.
- Analyze the buffering capacity of different buffer solutions.
- Interpret the behavior of some metals against a solution of HCl, according to its reducing power. Check the reaction products.
- Obtain a metal from a solution of one of its salts by reacting it with another more reducing metal.
- Study the influence of some factors on redox reactions, such as pH or complex formation.
- Build galvanic batteries with the appropriate assembly (electrodes, salt bridge, voltmeter, etc.) and predict the theoretical voltage that the system should provide on the basis of reduction potentials.
- Prepare an assembly with an U-tube, to develop an electrolysis reaction from a potassium iodide solution. Identify the products formed on the electrodes with specific reactions.
- Prepare the appropriate assembly to analyze the migration of metal ions as an example of electrochemical process to solve environmental problems.
- Determine the hardness of a water sample by means of a complexometric titration using EDTA as

titrant and NET as indicator.
- Prepare the appropriate assembly to perform a softening process of a water sample by exchange of Ca^{2+} and Mg^{2+} ions for Na^{+}.
- Carry out a deionization process of a water sample, through exchange of cations and anions by H^{+} and OH^{-}, respectively.
- Carry out different tests to interpret whether the softening and deionization processes have been correct (ionic conductivity, pH, or chloride test).

DESCRIPTION OF CONTENTS

1. Seminar 1

Presentation.
Safety measures.
Materials and basic operations in the laboratory.
Waste minimisation program.

2. Practice 1: Intermolecular forces

Physical properties of chemical compounds.
Acid-base reactions and solubility.

3. Practice 2. Synthesis and Purification of an Organic Compound.

Synthesis, isolation and purification of Acetanilide.
Reflux system.
Purification by crystallisation.
Characterisation by melting point and thin-layer chromatography.

4. Practice 3. Thermochemistry.

Determination of the calorific capacity of a calorimeter. Determination of the enthalpy variation of a neutralization reaction. Estimation of the dissolution heat of an ionic solid. Effect of temperature on the solubility of an ionic solid.

5. Seminar 2

How to write a Laboratory Report.
Objectives, index and theoretical introduction.
Treatment and discussion of results.
Formal aspects. Presentation of Tables and Figures.
Bibliography.

6. Practice 4. Chemical equilibrium.

Chemical reactions in the test tube.
Factors influencing chemical equilibrium.
Reversible and irreversible reactions.

7. Practice 5. Colligative Properties.

Freezing-point depression. Cryoscopic constant.
Molality.
Determination of molecular weights by cryoscopy.

8. Practice 6. Kinetics (1).

Discoloration kinetics of the "violet crystal". Instantaneous rate constant. Experimental determination of the rate constant and the order of reaction. Photocolorimetric technique. Apparent rate constants and absolute rate constant.

9. Seminar 3

Analysis and discussion of the results of practice sessions P1 to P6.

10. Practice 7. Kinetics (2).

Kinetics of the decomposition reaction of hydrogen peroxide.
Use of catalyst (potassium iodide).
Factors affecting the rate of decomposition: reactive concentration, catalyst and temperature.
Rate law. Activation energy.

11. Practice 8. Potentiometric Titration.

Primary standards.
Titration curves.
Determination of Kw.
Determination of acidity constant of acetic acid.

12. Practice 9. Buffer Solutions of pH.

Study of the buffering capacity of buffer solutions.
Preparation of buffer solutions.
Effect of adding bases or acids to buffer solutions.
Buffering capacity.

13. Practice 10. Electrochemistry.

Behaviour of some metals in HCl solution.
Influence of pH and complex formation on redox reactions.
Construction of galvanic cells.
Electrolysis.

14. Practice 11. Determining Water Hardness.

Determination of the hardness of a water sample by means of a complexomètric titration with EDTA. Ionic exchange. Softening and deionization. Measurements of ionic conductivity and pH. Chlorides test.

15. Seminar 4

Analysis and discussion of the results of practice sessions P7 to P11.

16. Evaluation

Final evaluation session.

WORKLOAD

ACTIVITY	Hours	% To be attended
Laboratory practices	48,00	100
Theory classes	12,00	100
Development of individual work	20,00	0
Study and independent work	50,00	0
Preparation of evaluation activities	10,00	0
Preparation of practical classes and problem	10,00	0
TOTAL	**150,00**	

TEACHING METHODOLOGY

Among the training activities described for the subject "Chemistry" in the verification report of the Degree in Chemistry, in this subject two are used: practical laboratory classes and seminars.In the practical laboratory sessions, an overview of the basic work of a chemistry laboratory will be offered. It is intended that students continue to acquire skills in the execution of the basic techniques of laboratory work. We want to get them to apply everything developed in the Chemistry Laboratory I to specific experiments, as well as to introduce some techniques that were not seen in that laboratory (in the first semester). A standard session will consist in the initial discussion of the previous questions that each practice has (that the student must bring resolved), and that will serve as a base to introduce the theoretical concepts on which the practice is based and to discuss the possible doubts or special precautions that they are required. The important part of the session will be the work and handling of

materials and products, depending on the objectives of the practice (most of the experimental procedure must be recorded by the student in his laboratory notebook). And at the end of the session, it is convenient to share the results achieved, an interpretation of these results and a reflection on whether the proposed objectives have been achieved. Four additional independent seminars of the laboratory sessions have been programmed, which will serve to reinforce the learning of the same ones, either dealing with monographic topics (for example, requirements to adequately prepare the memory of a laboratory practice), either to solve or analyze doubts that have arisen in the treatment and interpretation of the results of the practices.

EVALUATION

Attendance at practical laboratory classes is mandatory. Justified absence will be allowed for a maximum of two sessions (preferably, it should be suggested to be recovered in some other subgroup).The assessment of student learning will be formative in nature and will be carried out by addressing different aspects that are part of two blocks with well differentiated characteristics:**a) Continuous evaluation** Those aspects that require a continuous evaluation of the progress and work developed throughout the course are part of this section. For this, the following will be taken into account: active participation in the seminars, the resolution of all those questions and problems that are proposed to them to work autonomously, and of course, the management in the laboratory, the monitoring of the security rules and the laboratory notebook.Since the work in the laboratory, the preparation work of the experience and the preparation of the notebook involves a continuous evaluation process throughout the course, the grade obtained for these three sections, in the first call, will be maintained in the second one. The sections listed below, together with the percentage of the grade, can not be recovered, if necessary, in the second call. Only in the case of the Notebook will a partial recovery of those sections corresponding to the treatment and interpretation of the results be allowed. i. Preparation of the experience (including the preliminary questions): 20% ii. Work in the Laboratory: 20% iii. Laboratory notebook: 20%**b) Evaluation of specific Activities** The acquired knowledge and skills will be evaluated through exams throughout the course. The presentation, oral and written, of a laboratory report is also part of this section. iv. Memory of a laboratory practice: 20% v. Evaluation exercises: 20%To be able to pass the subject, a grade equal to or higher than 4 points is required in each of the five sections that make up the evaluation, and the weighted sum of all of them will reach 5 points. **Certificate of Honor Test**For those students who, applying the above evaluation criteria, obtain a final mark of excellence and the teacher considers him worthy of the MH qualification, they will have the possibility to perform a joint additional test for all students of all the groups enrolled in the subject. It will consist of a test with short questions and/or test type that will cover all the practices carried out.

In any case, the evaluation system will be governed by the provisions of the *Evaluation and Qualification Regulations of the University of Valencia for Degrees and Masters*

(http://www.uv.es/graus/normatives/2017_108_Reglament_avaluacio_qualificacio.pdf)

REFERENCES

Basic

- Petrucci, R.H.; Herring, F.G.; Madura, J.D.; Bissonnette, C. Química General. Principios y aplicaciones modernas, 10ª edición. Madrid, Pearson Educación, 2011
ISBN: 978-84-8322-680-3 (CI 54 PET)
- Chang, R. y Goldsby, K.A. Química, 11ª edición, México. Ed McGraw Hill, 2013
ISBN: 978-607-15-09284 (CI 54 CHA)
- González, R. Química General para las Ciencias Ambientales, Publicacions de la Universitat de València, (2011) ISBN: 9788437081700 (CI 628 GON)
- Olba, A. Quimica General. Equilibri i canvi València, Universitat de València, Servei de Publicacions, 2007. ISBN: 978-8437068435 (CI 54 OLB)
- Petrucci, R.H. et al. 11ª edición, 2017 (on-line)
http://www.ingebook.com/ib/NPcd/IB_BooksVis?cod_primaria=1000187&codigo_libro=6751
- Chang, R.; Goldsby, K.A., 11ª edición, 2013 (on-line)
http://www.ingebook.com/ib/NPcd/IB_BooksVis?cod_primaria=1000187&codigo_libro=4277

Additional

- Peterson, W.R. Introducción a la nomenclatura de sustancias química, Barcelona, Edit. Reverté, 2010
ISBN: 978-84-29175721
- Brown, T.L. et al. Química. La Ciencia Central, 12ª edición. México, Pearson Educación, 2013 ISBN: 978-607-32-2237-2 (CI 54 QUI)
- Brown, T.L. et al. , 12ª edición, 2014 (on-line)
http://www.ingebook.com/ib/NPcd/IB_BooksVis?cod_primaria=1000187&codigo_libro=4690