

**COURSE DATA****Data Subject**

Code	34185
Name	Chemistry laboratory I
Cycle	Grade
ECTS Credits	6.0
Academic year	2018 - 2019

Study (s)

Degree	Center	Acad. Period
1110 - Degree in Chemistry	Faculty of Chemistry	1 First term

Subject-matter

Degree	Subject-matter	Character
1110 - Degree in Chemistry	1 - Chemistry	Basic Training

Coordination

Name	Department
MILIAN MEDINA, BEGOÑA	315 - Physical Chemistry

SUMMARY

This subject is compulsory and of basic character, is taught in the first semester of the first year of the Degree in Chemistry, with a volume of 6 credits. Together with the "Laboratory of Chemistry II" (also obligatory of basic character, but taught in the second semester), it is intended, essentially, that the student learn the operation and the basic work techniques that will be developed in a laboratory chemical; and the preparation, recording, analysis and presentation of results of an experimental work. In this way, the essential foundations will be established so that the experiences of the different branches that make up the discipline can subsequently be successfully addressed.

In this specific subject, the security, analysis and interpretation of data necessary for the development of any chemical experience, as well as the management and treatment of data developed in any chemical laboratory will be addressed. To this end, experiments will be carried out in which different basic techniques must be used, so that later they can be applied to more complex tests.



It is assumed that students know and use, in a basic but clear way, the concepts taught in the last year of High School Chemistry. However, all the scripts include a theoretical introduction and, whenever necessary, additional teaching material will be provided to cover those deficiencies that are detected.

PREVIOUS KNOWLEDGE

Relationship to other subjects of the same degree

There are no specified enrollment restrictions with other subjects of the curriculum.

Other requirements

Students are expected to know and to be able to clearly apply the concepts taught on their last High School Chemistry course, albeit at a basic level. However, all scripts will include a theoretical introduction and, when required, extra teaching materials will be provided.

OUTCOMES

1108 - Degree in Chemistry

- Develop capacity for analysis, synthesis and critical thinking.
- Show inductive and deductive reasoning ability.
- Demonstrate leadership and management skills, entrepreneurship, initiative, creativity, organization, planning, control, leadership, decision making and negotiation.
- Demonstrate ability to work in teams both in interdisciplinary teams and in an international context.
- Demonstrate ability to communicate information, ideas, problems and solutions to both specialist and non-specialist audiences and using information technology, as appropriate.
- Demonstrate a commitment to ethics, equality values and social responsibility as a citizen and as a professional.
- Acquire a permanent sensitivity to quality, the environment, sustainable development and the prevention of occupational hazards.
- Demonstrate knowledge of the main aspects of chemical terminology, nomenclature, conventions and units.
- Interpret the variation of the characteristic properties of chemical elements according to the periodic table.
- Demonstrate knowledge of the characteristics and behaviour of the different states of matter and the theories used to describe them.



- Demonstrate knowledge of the principles of thermodynamics and kinetics and their applications in chemistry.
- Relate the macroscopic properties and the properties of individual atoms and molecules, including macromolecules (natural and synthetic), polymers, colloids and other materials.
- Demonstrate knowledge and understanding of essential facts, concepts, principles and theories related to the areas of chemistry.
- Solve qualitative and quantitative problems following previously developed models.
- Evaluate, interpret and synthesise chemical data and information.
- Handle chemicals safely.
- Carry out standard experimental procedures involved in synthetic and analytical work, in relation to organic and inorganic systems.
- Interpret data from observations and measurements in the laboratory in terms of their significance and the theories that underpin them.
- Evaluate the risks in the use of chemicals and laboratory procedures.
- Recognise and evaluate chemical processes in daily life.
- Understand the qualitative and quantitative aspects of chemical problems.
- Students must have acquired knowledge and understanding in a specific field of study, on the basis of general secondary education and at a level that includes mainly knowledge drawn from advanced textbooks, but also some cutting-edge knowledge in their field of study.
- Students must be able to apply their knowledge to their work or vocation in a professional manner and have acquired the competences required for the preparation and defence of arguments and for problem solving in their field of study.
- Students must have the ability to gather and interpret relevant data (usually in their field of study) to make judgements that take relevant social, scientific or ethical issues into consideration.
- Students must be able to communicate information, ideas, problems and solutions to both expert and lay audiences.
- Students must have developed the learning skills needed to undertake further study with a high degree of autonomy.
- Express oneself correctly, both orally and in writing, in any of the official languages of the Valencian Community.
- Have basic skills in the use of information and communication technology and properly manage the information obtained.

LEARNING OUTCOMES



At the end of the subject "Laboratory of Chemistry I", the student must be able to:

- Distinguish and recognize the most commonly used laboratory material: glass (volumetric and non-volumetric), electrical material and assemblies (heating mantle, distillation, rotary evaporator, scales, melting point apparatus, etc.).
- Know the waste minimization protocols.
- Understand and distinguish the labelling information of laboratory products, especially that referring to Safety Standards, H and P phrases, pictograms, etc.
- Use common appliances such as the Bunsen burner or the vacuum pump.
- Distinguish between the different types of filtration, depending on the objective pursued.
- Precisely prepare a conical filter and a pleat filter.
- Separate well differentiated products according to their solubility, using processes such as decanting and choosing the right solvent.
- Handle with precision the different types of filtering, both hot and cold.
- Estimate the amount of solute that can be dissolved in a solvent depending on its solubility (obtained from the literature).
- Know the phase changes that can occur in a compound when we are in the laboratory.
- Know the safety measures to take into account in the handling and heating of flammable liquids.
- Mount a simple distillation equipment whose objective is to measure the boiling point of a liquid.
- Calculate the efficiency (yield) of the distillation process.
- Properly measure the melting point of a crystallized solid.
- Handle the melting point apparatus correctly and select the appropriate program to the necessary extent.
- Know the liquid-liquid extraction technique.
- Know what characteristics an organic solvent must possess to use it in an extraction.
- Handle with precision a separating funnel, taking into account the safety protocol.
- Know the procedure to isolate the aqueous phase and the organic phase of the extraction process.
- Know at what time and to which phase a desiccant should be added (anhydrous salt).
- Use the magnitudes and their units correctly in the laboratory processes that involve measurements or quantitative calculations.
- Estimate adequately the errors made in the measurements (absolute error, relative error, standard deviation, etc.).
- Correctly perform the hot filtration process to achieve the most perfect crystallization possible.
- Use the thin layer chromatography technique to identify a previously purified compound.
- Choose the appropriate eluent according to the polarity of the compound to be identified.
- Distinguish adequately between the functions of the stationary phase and the eluent in the thin layer chromatography technique.
- Know which are the possible eluents to use and know how to order them by their polarity.
- Accurately calculate the amount of solid or liquid necessary to prepare a solution of determined concentration.
- Accurately handle the volumetric material in the process of preparing a solution.
- Qualitatively and quantitatively determine the expected pH value for prepared solutions (both acids and solid salts).
- Accurately use the pH meter in the process of measuring the pH of a solution.
- Know the procedure of standardization of a solution and the volumetric material necessary for it.
- Manipulate properly a burette to make an assessment.
- Know what an indicator is and what are the conditions in which its use is useful: in what interval it turns and for what type of valuations it is adequate.
- Know the use of primary standards and their characteristics.



- Determine the concentration of a solution from a valuation process, calculating the errors committed, the standard deviation, etc.
- Precisely prepare solutions by dilution, from a stock solution.
- Know what it is and how a white solution is used.
- Know the use of a simple spectrophotometer for the measurement of the absorbance spectrum of a coloured solution (such as CuSO_4).
- Determine the analytical wavelength for a problem solution.
- Apply the Law of Lambert-Beer to relate the absorbance with the concentration of a concrete solution.
- Draw accurately the calibration line that relates the absorbance to the concentration of a solution.
- Determine, from the calibration line, the concentration of a problem solution.
- Prepare an assembly to carry out a distillation, both simple and with fractionation column.
- Apply the protocol correctly to properly separate, by distillation, two miscible liquids.
- Calculate the concentration of acid in the distilled phase of a mixture with acetic acid and acetone; and express it in different habitual units (mol / L , g / L , mole fraction).
- Analyze the necessary parameters to compare the separation of two miscible liquids by simple and fractional distillation.
- Determine the density of a mixture of two known liquids by weighing.
- Know the use of a eudiometer as a precision glass material to collect the gas generated in a reaction, and how to prepare the assembly.
- Do stoichiometric calculations applied to a reaction where there is a limiting reagent.
- Determine the molar mass of CaCO_3 by two methods: gravimetric and volumetric.
- Estimate the weight gain of CaCO_3 in a problem sample.

DESCRIPTION OF CONTENTS

0. Prevention Session

Prevention and fire performance in buildings for teaching-university use.

1. Seminar 1

Presentation.

Management and organisation of laboratory work.

Preparation of experimental work.

2. Practice 1. Safety and Laboratory Material.

Safety rules. Simplified sheets of compounds. Pictograms. H and P phrases. Laboratory material (glass material, electrical material, assemblies, lighter, vacuum pump, etc.). Types of filtration. Use of the balance. Direct weighing and tare. Waste. Waste minimization program.



3. Practice 2. Dissolution, precipitation and crystallization.

Dissolution and solubility. Precipitation and crystallization.
Solid-liquid separations: decantation and filtration.

4. Practice 3. Characterisation of Liquids and Solids.

Distillation. Boiling point determination.
Melting point determination.

5. Seminar 2

Presentation of results.
Physical magnitudes. Units system.
Measurement and experimental error.
Accuracy and precision. Significant figures.

6. Practice 4. Liquid-Liquid Extraction

Separation and isolation of organic unknown compounds.
Extracting solvents.
Aqueous phase and organic phase.

7. Practice 5. Crystallization and identification of compounds.

Session A: Purification (crystallization) and identification of an organic acid.
Session B: Purification (crystallization) and identification of a neutral compound.
Characterisation and identification by melting point.
Thin layer chromatography.

8. Seminar 3

Analysis and discussion of the results of practices P2 to P5.

9. Practice 6. Preparation of Solutions and Measurement of pH.

Acidity, alkalinity, equilibrium and pH.
Preparing solutions of different concentrations.
Solutions from commercial products (solid salts).
Use of the pH-meter and pH measurements.



10. Practice 7. Acid-Base Titration.

Stoichiometry and neutralisation of acid-base reactions.
Indicators in acid-base titrations.
Use of primary and secondary standards.

11. Practice 8. Absorbance spectrum of solutions.

Aqueous solutions of CuSO_4 by dilution.
Preparation and utility of a blank solution.
Use of the visible spectrophotometer and spectrum plot.
Absorbance measurements of copper sulphate solutions.
Data treatment.

12. Practice 9. Distillation of Mixtures of Miscible Liquids

Distillation acetone-acetic acid.
Simple distillation and distillation with a fractionating column. Effectiveness of both processes.
Mixture density by weighing.

13. Practice 10. Stoichiometric Calculations.

Reaction between calcium carbonate and hydrochloric acid.
Molar mass determination of CaCO_3 .
Percentage purity of an unknown sample.
Gravimetric and volumetric methods.

14. Seminar 4

Analysis and discussion of the results of practices P6 to P10.

15. Evaluation

Final evaluation session.

**WORKLOAD**

ACTIVITY	Hours	% To be attended
Laboratory practices	48,00	100
Tutorials	12,00	100
Development of individual work	20,00	0
Study and independent work	50,00	0
Preparation of evaluation activities	10,00	0
Preparation of practical classes and problem	10,00	0
TOTAL	150,00	

TEACHING METHODOLOGY

Among the training activities described for the subject "Chemistry" in the verification report of the Degree in Chemistry, in this subject two are used: practical laboratory classes and seminars.

In the practical laboratory sessions, an overview of the basic work of a chemistry laboratory will be offered. It is intended that students acquire skills in the execution of the basic techniques of laboratory work. They should become familiar with the mechanisms of safety and management, handling of material and equipment, treatment and presentation of data, decision making and choosing the most appropriate procedure, if applicable. A standard session will consist in the initial discussion of the previous questions that each practice has (that the student must bring resolved), and that will serve as a base to introduce the theoretical concepts on which the practice is based and to discuss the possible doubts or special precautions that they are required. The important part of the session will be the work and handling of materials and products, depending on the objectives of the practice (most of the experimental procedure must be recorded by the student in his laboratory notebook). And at the end of the session, it is convenient to share the results achieved, an interpretation of these results and a reflection on whether the proposed objectives have been achieved.

Four additional independent seminars of the laboratory sessions have been programmed, which will serve to reinforce the learning of the same ones, either treating monographic subjects (for example, treatment of magnitudes, units and calculation of errors), either to solve or analyze doubts that have arisen in the treatment and interpretation of the results of the practices.

Since it is the first laboratory that first-year students have access to, two additional activities related to waste prevention and management are planned:



- Workshop on Prevention and extinction of fires, given by the chief prevention officer of the Valencia Provincial Fire Department Consortium.

- Conference on waste treatment in the laboratories of the Faculty of Chemistry, taught by a technician of the General Chemistry Laboratory, and whose objective is to make students aware of the process of minimization and correct waste management of a laboratory of these characteristics.

EVALUATION

Attendance at practical laboratory classes is mandatory. Justified absence will be allowed for a maximum of two sessions (preferably, it should be suggested to be recovered in some other subgroup).

The assessment of student learning will be formative in nature and will be carried out by addressing different aspects that are part of two blocks with well differentiated characteristics:

a) Continuous evaluation

Those aspects that require a continuous evaluation of the progress and work developed throughout the course are part of this section. For this, the following will be taken into account: active participation in the seminars, the resolution of all those questions and problems that are proposed to them to work autonomously, and of course, the management in the laboratory, the monitoring of the security rules and the laboratory notebook.

Since the work in the laboratory, the preparation work of the experience and the preparation of the notebook involves a continuous evaluation process throughout the course, the grade obtained for these three sections, in the first call, will be maintained in the second one. The sections listed below, together with the percentage of the grade, can not be recovered, if necessary, in the second call. Only in the case of the Notebook will a partial recovery of those sections corresponding to the treatment and interpretation of the results be allowed.

- Preparation of experience and work in the laboratory (20%)
- Deliverables (previous, post, results) (30%)
- Laboratory notebook (20%)

In total, this section: **70%** of the final grade

b) Evaluation of specific Activities

The acquired knowledge and skills will be evaluated through exams throughout the course. Any questionnaire or activity carried out in the seminar sessions is also part of this section.



Evaluation exercises: **30%** of the final grade

To be able to pass the subject, a grade equal to or greater than 4 points is required in each of the two blocks that make up the evaluation, and the weighted sum of both will reach 5 points.

Certificate of Honor Test

For those students who, applying the above evaluation criteria, obtain a final mark of excellence and the teacher considers him worthy of the MH qualification, they will have the possibility to perform a joint additional test for all students of all the groups enrolled in the subject. It will consist of a test with short questions and/or test type that will cover all the practices carried out.

In any case, the evaluation system will be governed by the provisions of the *Evaluation and Qualification Regulations of the University of Valencia for Degrees and Masters*

http://www.uv.es/graus/normatives/2017_108_Reglament_avaluacio_qualificacio.pdf

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